

STEPS:

- 1) Count up valence electrons of all atoms in molecule. (Make sure you have all NONMETALS.)
- 2) Locate central atom (least electronegative usually).
- 3) Arrange other atoms around central atoms with single bonds.
- 4) Use lone pairs to fill in any incomplete octets. (check to make sure each atom has exactly an octet)
- 5) Count up electrons in drawn structure, compare to original amount of valence electrons available.
- 6) If too many electrons are in drawn structure, try double or triple bonds until electron number matches.

Formula: Br₂
VE:

Formula: HCl
VE:

Formula: CBr₄
VE:

Formula: SeF₂
VE:

Formula: F₂Se
#VE:

Formula: SiI₄
#VE:

Formula: NF₃
#VE:

Formula: Si₂H₂
#VE:

Formula: C₂H₄
#VE:

Formula: SeO₂
#VE:

Formula: CH₂O
#VE:

Formula: C₂H₂
#VE:

Formula: N₂
#VE:

Formula: FCN
#VE:

Formula: H₂
#VE: