**14.3 Guided Reading**

**Integrated Science – Matter Name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_Per:\_\_\_**

1. How is energy involved in chemical reactions?
2. Describe and give an example of an exothermic reaction.
3. Describe and give an example of an endothermic reaction.
4. Activation energy is the energy needed to begin a reaction and \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ the chemical bonds.
5. What kind of reaction is this (endothermic or exothermic)? How do you know?

2 Al2O3 + *Energy*  🡪 4 Al + 3 O2

1. A \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_reaction is a type of endothermic reaction that takes place when an ionic compound mixes in water to create an ionic solution.
2. The change in concentration of reactants and products in a chemical reaction, occurring over a period of time is called \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
3. Substances added to a reaction that either increase or inhibit the rate of the reaction are \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ and \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.