**13.2 Guided Reading**

**Integrated Science – Matter Name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_Per:\_\_\_**

1. All compounds have an electrical charge of \_\_\_\_\_\_\_\_\_\_\_\_\_\_.
2. An oxidation number is the quantity that indicates the charge on an atom when it has gained, lost, or \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ electrons.
3. Fill in the following table:

|  |  |  |
| --- | --- | --- |
| **Atom** | **e- lost or gained** | **Oxidation #** |
| Na |  |  |
| Ba |  |  |
| Al |  |  |
| P |  |  |
| Cl |  |  |

1. Would Beryllium tend to lose two electrons or gain six when forming bonds?
2. What is the most common oxidation number for group 13 on the Periodic Table?
3. The farther apart elements are on the Periodic Table the more likely they are to form \_\_\_\_\_\_\_\_\_\_ bonds.
4. Nonmetals tend to form \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_bonds with each other.
5. How many chlorine atoms are needed to bond with a calcium atom to form a compound?
6. What is a binary compound?
7. How many atoms of each element is in MgCO3?
8. An ion that is composed of more than one atom is called a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ ion.
9. What is the oxidation number for nitrate?