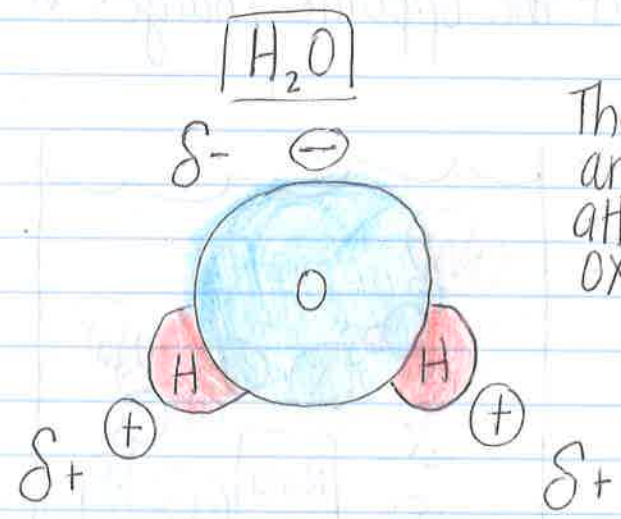


Polar & Non-Polar Compounds

Electronegativity (EN)

The attraction of a particular atom for the electrons of a bond.

The greater the EN, the more attracted the electrons are to the nucleus of the atom.

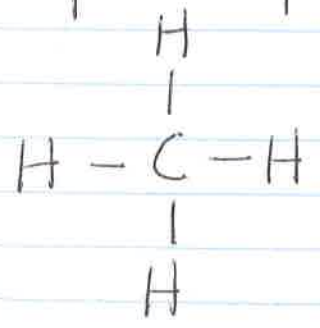


The electrons are more attracted to the oxygen atom.

Water is a polar compound. The electrons are not shared equally, so each end of the molecule has a + or - charge. However, the compound as a whole is neutral.

Like dissolves like. Polar dissolves polar.

Non-polar compound: The electrons are shared equally.



Non-polar compounds such as oils and fats do not dissolve in polar compounds such as water.

Water is an excellent solvent. It can pull apart ionic & polar substances in solution.

When NaCl or salt is placed in water, the \oplus & \ominus ends of the H_2O molecule will attract the opposite charges or ions.

