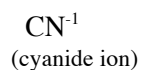
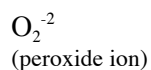
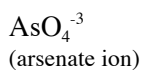
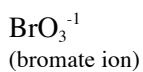
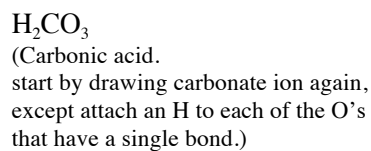
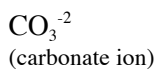
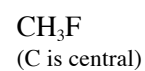
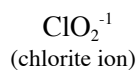
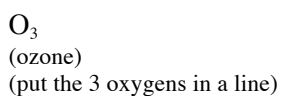
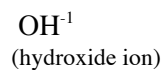


Write the total number of valence electrons next to each formula, and then draw the Lewis dot structure for each atom or ion.



Reactions Review!

Write a balanced chemical equation for each reaction, including phase subscripts on all reactants and products.

a. Bromine reacts with aluminum. _____

b. Calcium burns. _____

c. Liquid butane (C₄H₁₀) is combusted.

2. Complete and balance each reaction. Show phase subscripts on all reactants and products.

a. S_(s) + K_(s) -----> _____

b. K_(s) + N_{2(g)} -----> _____

c. Chlorine reacts with iron. Assume that the iron forms *ferric* ion.

d. C₁₀H₁₈O_(l) + O_{2(g)} -----> _____

e. C₆H₁₂O_{3(l)} + O_{2(g)} -----> _____

f. O_{2()} + C_() -----> _____

g. O_{2()} + H_{2()} -----> _____

h. O_{2()} + Li_() -----> _____